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Chapter 14. Chemical Kinetics

14.4. Use the equations in the AP Chemistry test booklet to work kinetics problems. Explain the concept of reaction half-life and describe the relationship between half-life and rate constant for a first-order reaction. Use graphical analysis to determine whether the rate law for a reaction is first or second order. 14.5-14.7

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14 Chemical Kinetics Reaction Rates • All reactions slow down over time. • Therefore, the best indicator of the rate of a reaction is the instantaneous rate near the beginning. C4H9Cl (aq) + H 2O(l) ? C4H9OH (aq) + HCl (aq) PDF Created with deskPDF PDF Writer - Trial ::

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14.1: Factors that Affect Reaction Rates. chemical kinetics - area of chemistry dealing with speeds/rates of reactions. rates of reactions affected by four factors. concentrations of reactants. temperature at which reaction occurs. presence of a catalyst. surface area of solid or liquid reactants and/or catalysts.

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The slope of this ... A.P. Chemistry Practice Test: Ch. 12, Kinetics MULTIPLE ... Major topics: integrated zero/first/second order rate laws, zero/first/second order reactions graphically, & half-life Page 14/22

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It has been found that for a chemical reaction with rise in temperature by 10 °C, the rate constant gets nearly doubled. 15. The temperature coefficient of a reaction is the ratio of the rate constants of the reaction at two temperatures differing from one another by 10°C.

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