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3-1 Calculate the specific volume of an air-
vapor mixture in cubic meters per kilogram
of dry air when the following conditions
prevail: $t = 30\text{ C}$, $W = 0.015\text{ kg/kg}$, and $p_t = 90\text{ kPa}$.
Solution: Equation 3-4. $T = 30\text{ C} + 273 = 303\text{ K}$
 $R_a = 287\text{ J/kg.K}$ $p_t = 90\text{ kPa} = 90,000\text{ Pa}$
Equation 3-2 $v_t = \frac{R_a T}{p_t} \frac{1}{1 - W} = \frac{287 \times 303}{90,000} \frac{1}{1 - 0.015} = 1.35 - 0.15p_s$

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Solution: $q = mc_p (t_2 - t_1)$ $m = 2.5\text{ kg/s}$ $c_p = 1.0\text{ kJ/kg.K}$ $t_2 = 30\text{ C}$ $t_1 = -10\text{ C}$

CHAPTER 2 - THERMAL PRINCIPLES

Page 2 of 5 Then, $q = (2.5) (1.0) (30 + 10)$ $q = 100\text{ kw}$ - - - Ans. 2-3.

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